

**GOVERNMENT COLLEGE FOR WOMEN,
PARADE GROUND, JAMMU
(An Autonomous College)**

CHEMISTRY SYLLABUS

B.Sc. SEMESTER I-VI

UNDER

CHOICE BASED CREDIT SYSTEM (CBCS)

GOVERNMENT COLLEGE FOR WOMEN, PARADE GROUND, JAMMU
(An Autonomous College)

Syllabi and Courses of study in Chemistry for B.Sc. Semester I-VI (CBCS)

Core Courses (C):

Semester	Course Number	Course Title	Credits
I	UCHTC 101	Atomic Structure, Bonding, General Organic Chemistry & Aliphatic Hydrocarbons	4
	UCHPC 101	Laboratory Course: Chemistry-I	2
II	UCHTC 201	Chemical Energetics, Equilibria & Functional Organic Chemistry	4
	UCHPC 201	Laboratory Course: Chemistry-II	2
III	UCHTC 301	Solutions, Phase Equilibrium, Conductance, Electrochemistry & Functional Group Organic Chemistry	4
	UCHPC 301	Laboratory Course: Chemistry-III	2
IV	UCHTC 401	Coordination Chemistry, States Of Matter & Chemical Kinetics	4
	UCHPC 401	Laboratory Course: Chemistry-IV	2

Skill Enhancement Courses (SEC):

Semester	Course Number	Course Title	Credits
III	UCHTS 301	Fuel Chemistry (Skill Enhancement Course SEC-1)	4
IV	UCHTS 401	Pesticide Chemistry (Skill Enhancement Course SEC-2)	4
V	UCHTS 501	Cosmetics, Perfumes & Medicinal Agents from Natural Sources (Skill Enhancement Course SEC-3)	4
VI	UCHTS 601	Green Methods In Chemistry (Skill Enhancement Course SEC-4)	4

Discipline Specific Elective Courses (DSE):

Semester	Course Number	Course Title	Credits
V	UCHTDSE 501	Spectroscopy, Photochemistry, Organometallics & Bioinorganic Chemistry (DSE 1)	4
	UCHPDSE 501	laboratory Course: Chemistry - V	2
V	UCHTDSE 511	Quantum Mechanics, Polymeric Materials, Industrial Metallurgy & Environment (DSE 2)	4
	UCHPDSE 511	laboratory Course: Chemistry - V	2
VI	UCHTDSE 601	Inorganic Materials of Industrial Importance & Organic Spectroscopy (DSE 3)	4
	UCHPDSE 601	laboratory Course: Chemistry - VI	2
VI	UCHTDSE 611	Analytical Methods, Electronic and Magnetic Properties of Transition Metals & Synthetic Organic Chemistry (DSE 4)	4
	UCHPDSE 611	laboratory Course: Chemistry - VI	2

Note:- Select two Papers, one each in Semester V (DSE 1 or DSE 2) & Semester VI (DSE 3 or DSE 4)

GOVERNMENT COLLEGE FOR WOMEN, PARADE GROUND, JAMMU
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**Syllabi and courses of study in chemistry for B.Sc. Semester I for the examinations to
be held in Dec. 2017 & 2018**

Course No.: UCHTC 101

Theory: 60 Lectures

**Title: ATOMIC STRUCTURE, BONDING, GENERAL ORGANIC CHEMISTRY &
ALIPHATIC HYDROCARBONS**

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Inorganic Chemistry-I (30 Periods)

Atomic Structure

(14 Lectures)

What is Quantum mechanics? Time independent Schrodinger equation and meaning of various terms in it. Significance of ψ and ψ^2 , Schrödinger equation for hydrogen atom. Radial and angular parts of the hydrogenic wavefunctions (atomic orbitals) and their variations for $1s$, $2s$, $2p$, $3s$, $3p$ and $3d$ orbitals (Only graphical representation). Radial and angular nodes and their significance. Radial distribution functions and the concept of the most probable distance with special reference to $1s$ and $2s$ atomic orbitals. Significance of quantum numbers, orbital angular momentum and quantum numbers m_l and m_s . Shapes of s , p and d atomic orbitals, nodal planes.

Rules for filling electrons in various orbitals, Electronic configurations of the atoms. Stability of half-filled and completely filled orbitals, concept of exchange energy. Relative energies of atomic orbitals, Anomalous electronic configurations.

Chemical Bonding and Molecular Structure

(16 Lectures)

Ionic Bonding: General characteristics of ionic bonding. Energy considerations in ionic bonding, lattice energy and solvation energy and their importance in the context of stability and solubility of ionic compounds. Statement of Born-Landé equation for calculation of lattice energy, Born-Haber cycle and its applications, polarizing power and polarizability. Fajan's rules, ionic character in covalent compounds, bond moment, dipole moment and percentage ionic character.

Covalent bonding: VB Approach: Shapes of some inorganic molecules and ions on the basis of VSEPR and hybridization with suitable examples of linear, trigonal planar, square planar, tetrahedral, trigonal bipyramidal and octahedral arrangements.

Concept of resonance and resonating structures in various inorganic compounds.

MO Approach: Rules for the LCAO method, bonding and antibonding MOs and their characteristics for *s-s*, *s-p* and *p-p* combinations of atomic orbitals, nonbonding combination of orbitals, MO treatment of homonuclear diatomic molecules of 1st and 2nd periods (including idea of *s-p* mixing) and heteronuclear diatomic molecules such as CO, NO and NO⁺. Comparison of VB and MO approaches.

Section B: Organic Chemistry-I (30 Periods)

Fundamentals of Organic Chemistry

(8 Lectures)

Physical Effects, Electronic Displacements: Inductive Effect, Electromeric Effect, Resonance and Hyperconjugation. Cleavage of Bonds: Homolysis and Heterolysis.

Structure, shape and reactivity of organic molecules: Nucleophiles and electrophiles. Reactive Intermediates: Carbocations, Carbanions and free radicals.

Strength of organic acids and bases: Comparative study with emphasis on factors affecting pK values. Aromaticity: Benzenoids and Hückel's rule.

Stereochemistry

(10 Lectures)

Conformations with respect to ethane, butane and cyclohexane. Interconversion of Wedge Formula, Newmann, Sawhorse and Fischer representations. Concept of chirality (upto two carbon atoms). Configuration: Geometrical and Optical isomerism; Enantiomerism, Diastereomerism and Meso compounds). Threo and erythro; D and L; *cis-trans* nomenclature; CIP Rules: R/ S (for upto 2 chiral carbon atoms) and E / Z Nomenclature (for upto two C=C systems).

Aliphatic Hydrocarbons

(12 Lectures)

Functional group approach for the following reactions (preparations & reactions) to be studied in context to their structure.

Alkanes: (Upto 5 Carbons). *Preparation:* Catalytic hydrogenation, Wurtz reaction, Kolbe's synthesis, from Grignard reagent. *Reactions:* Free radical Substitution: Halogenation.

Alkenes: (Upto 5 Carbons) *Preparation:* Elimination reactions: Dehydration of alkenes and dehydrohalogenation of alkyl halides (Saytzeff's rule); *cis* alkenes (Partial catalytic hydrogenation) and *trans* alkenes (Birch reduction). *Reactions:* *cis*-addition (alk. KMnO₄) and *trans*-addition (bromine), Addition of HX (Markownikoff's and anti-Markownikoff's addition), Hydration, Ozonolysis, oxymercuration-demercuration, Hydroboration-oxidation.

Alkynes: (Upto 5 Carbons) *Preparation:* Acetylene from CaC_2 and conversion into higher alkynes; by dehalogenation of tetra halides and dehydrohalogenation of vicinal-dihalides. *Reactions:* formation of metal acetylides, addition of bromine and alkaline KMnO_4 , ozonolysis and oxidation with hot alkaline KMnO_4 .

Reference Books:

- Lee, J.D. *Concise Inorganic Chemistry* ELBS, 1991.
- Cotton, F.A., Wilkinson, G. & Gaus, P.L. *Basic Inorganic Chemistry*, 3rd ed., Wiley.
- Douglas, B.E., McDaniel, D.H. & Alexander, J.J. *Concepts and Models in Inorganic Chemistry*, John Wiley & Sons.
- Huheey, J.E., Keiter, E.A., Keiter, R.L. & Medhi, O.K. *Inorganic Chemistry: Principles of Structure and Reactivity*, Pearson Education India, 2006.
- Graham Solomon, T.W., Fryhle, C.B. & Snyder, S.A. *Organic Chemistry*, John Wiley & Sons (2014).
- McMurry, J.E. *Fundamentals of Organic Chemistry*, 7th Ed. Cengage Learning India Edition, 2013.
- Sykes, P. *A Guidebook to Mechanism in Organic Chemistry*, Orient Longman, New Delhi (1988).
- Eliel, E.L. *Stereochemistry of Carbon Compounds*, Tata McGraw Hill education, 2000.
- Finar, I.L. *Organic Chemistry* (Vol. I & II), E.L.B.S.
- Morrison, R.T. & Boyd, R.N. *Organic Chemistry*, Pearson, 2010.
- Bahl, A. & Bahl, B.S. *Advanced Organic Chemistry*, S. Chand, 2010.

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
- Weightage for attendance – 5 marks

External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

The question paper will have three sections.

Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

Section B will consist of five medium answer type questions (with explanation having 250 to 300 words) of seven marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 35 marks).

Section C will consist of five long answer type questions (with detailed explanation having 500 to 600 words) of fifteen marks each, covering the entire syllabus. The candidate shall have to attempt any two questions (Total marks: 30).

SEMESTER-I

Course No.: UCHPC 101 (Practicals)

(For examinations to be held in Dec. 2017 & 2018)

Title:- Laboratory Course: Chemistry-I

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Inorganic Chemistry - Volumetric Analysis

1. Estimation of sodium carbonate and sodium hydrogen carbonate present in a mixture.
2. Estimation of oxalic acid by titrating it with KMnO_4 .
3. Estimation of water of crystallization in Mohr's salt by titrating with KMnO_4 .
4. Estimation of Fe (II) ions by titrating it with $\text{K}_2\text{Cr}_2\text{O}_7$ using internal indicator.
5. Estimation of Cu (II) ions iodometrically using $\text{Na}_2\text{S}_2\text{O}_3$.

Section B: Organic Chemistry

1. Detection of extra elements (N, S, Cl, Br, I) in organic compounds (containing upto two extra elements)
2. Separation of mixtures by Chromatography: Measure the R_f value in each case (combination of two compounds to be given)
 - (a) Identify and separate the components of a given mixture of 2 amino acids (glycine, aspartic acid, glutamic acid, tyrosine or any other amino acid) by paper chromatography
 - (b) Identify and separate the sugars present in the given mixture by paper chromatography.

Reference Books:

- Svehla, G. *Vogel's Qualitative Inorganic Analysis*, Pearson Education, 2012.
- Mendham, J. *Vogel's Quantitative Chemical Analysis*, Pearson, 2009.
- Vogel, A.I., Tatchell, A.R., Furnis, B.S., Hannaford, A.J. & Smith, P.W.G., *Textbook of Practical Organic Chemistry*, Prentice-Hall, 5th edition, 1996.
- Mann, F.G. & Saunders, B.C. *Practical Organic Chemistry* Orient-Longman, 1960.

DISTRIBUTION OF MARKS

A.	Daily evaluation of practical records/viva-voce/attendance, etc.	Attendance: 05 marks
		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks (Two practicals of 10 marks each)
		Viva-voce: 05 marks

GOVERNMENT COLLEGE FOR WOMEN, PARADE GROUND, JAMMU
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**Syllabi and courses of study in chemistry for B.Sc. Semester II for the examinations to
be held in May 2018 & 2019**

Course No.: UCHTC 201

Theory: 60 Lectures

**Title: CHEMICAL ENERGETICS, EQUILIBRIA & FUNCTIONAL ORGANIC
CHEMISTRY**

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Physical Chemistry-I (30 Lectures)

Chemical Energetics

(10 Lectures)

Important principles and definitions of thermochemistry. Concept of standard state and standard enthalpies of formations, integral and differential enthalpies of solution and dilution. Calculation of bond energy, bond dissociation energy and resonance energy from thermochemical data. Variation of enthalpy of a reaction with temperature – Kirchhoff's equation.

Statement of Third Law of thermodynamics and calculation of absolute entropies of substances.

Chemical Equilibrium

(8 Lectures)

Free energy change in a chemical reaction. Thermodynamic derivation of the law of chemical equilibrium. Distinction between G and G° , Le Chatelier's principle. Relationships between K_p , K_c and K_x for reactions involving ideal gases.

Ionic Equilibria

(12 Lectures)

Strong, moderate and weak electrolytes, degree of ionization, factors affecting degree of ionization, ionization constant and ionic product of water. Ionization of weak acids and bases, pH scale, common ion effect. Salt hydrolysis-calculation of hydrolysis constant, degree of hydrolysis and pH for different salts. Buffer solutions. Solubility and solubility product of sparingly soluble salts – applications of solubility product principle.

Section B: Organic Chemistry-II (30 Lectures)

Functional group approach for the following reactions (preparations & reactions) to be studied in context to their structure:

Aromatic hydrocarbons (8 Lectures)

Preparation (Case benzene): from phenol, by decarboxylation, from acetylene, from benzene sulphonic acid.

Reactions: (Case benzene): Electrophilic substitution: nitration, halogenation and sulphonation. Friedel-Craft's reaction (alkylation and acylation) (upto 4 carbons on benzene). Side chain oxidation of alkyl benzenes (upto 4 carbons on benzene).

Alkyl and Aryl Halides (8 Lectures)

Alkyl Halides (Upto 5 Carbons) Types of Nucleophilic Substitution (SN^1 , SN^2 and SN^i) reactions.

Preparation: from alkenes and alcohols.

Reactions: hydrolysis, nitrite & nitro formation, nitrile & isonitrile formation. Williamson's ether synthesis: Elimination vs substitution.

Aryl Halides *Preparation*: (Chloro, bromo and iodo-benzene case): from phenol, Sandmeyer & Gattermann reactions.

Reactions (*Chlorobenzene*): Aromatic nucleophilic substitution (replacement by $-OH$ group) and effect of nitro substituent. Benzyne Mechanism: KNH_2/NH_3 (or $NaNH_2/NH_3$).

Reactivity and Relative strength of C-Halogen bond in alkyl, allyl, benzyl, vinyl and aryl halides.

Alcohols, Phenols and Ethers (Upto 5 Carbons) (14 Lectures)

Alcohols *Preparation*: Preparation of 1° , 2° and 3° alcohols: using Grignard reagent, Ester hydrolysis, Reduction of aldehydes, ketones, carboxylic acid and esters.

Reactions: With sodium, HX (Lucas test), esterification, oxidation (with PCC, alk. $KMnO_4$, acidic dichromate, conc. HNO_3). Oppeneauer oxidation.

Phenols: (Phenol case) *Preparation*: Cumene hydroperoxide method, from diazonium salts. *Reactions*: Electrophilic substitution: Nitration, halogenation and sulphonation. Reimer-Tiemann Reaction, Gattermann-Koch Reaction, Houben-Hoesch Condensation, Schotten – Baumann Reaction.

Ethers (aliphatic and aromatic): Cleavage of ethers with HI.

Aldehydes and ketones (aliphatic and aromatic): (Formaldehyde, acetaldehyde,

acetone and benzaldehyde)

Preparation: from acid chlorides and from nitriles.

Reactions – Reaction with HCN, ROH, NaHSO₃, NH₂-G derivatives. Iodoform test. Aldol Condensation, Cannizzaro's reaction, Wittig reaction, Benzoin condensation. Clemensen reduction and Wolf Kishner reduction. Meerwein-Ponndorf Verley reduction.

Reference Books:

- Graham Solomon, T.W., Fryhle, C.B. & Snyder, S.A. *Organic Chemistry*, John Wiley & Sons (2014).
- McMurry, J.E. *Fundamentals of Organic Chemistry*, 7th Ed. Cengage Learning India Edition, 2013.
- Sykes, P. *A Guidebook to Mechanism in Organic Chemistry*, Orient Longman, New Delhi (1988).
- Finar, I.L. *Organic Chemistry* (Vol. I & II), E.L.B.S.
- Morrison, R.T. & Boyd, R.N. *Organic Chemistry*, Pearson, 2010.
- Bahl, A. & Bahl, B.S. *Advanced Organic Chemistry*, S. Chand, 2010.
- Barrow, G.M. *Physical Chemistry* Tata McGraw-Hill (2007).
- Castellan, G.W. *Physical Chemistry* 4th Ed. Narosa (2004).
- Kotz, J.C., Treichel, P.M. & Townsend, J.R. *General Chemistry* Cengage Learning India Pvt. Ltd., New Delhi (2009).
- Mahan, B.H. *University Chemistry* 3rd Ed. Narosa (1998).
- Petrucci, R.H. *General Chemistry* 5th Ed. Macmillan Publishing Co.: New York (1985).

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Section B will consist of five medium answer type questions (with explanation having 250 to 300 words) of seven marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 35 marks).

Section C will consist of five long answer type questions (with detailed explanation having 500 to 600 words) of fifteen marks each, covering the entire syllabus. The candidate shall have to attempt any two questions (Total marks: 30).

SEMESTER-II

Course No.: UCHPC 201 (Practicals)

(For examinations to be held in May 2018 & 2019)

Title:- Laboratory Course: Chemistry-II

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Physical Chemistry

Thermochemistry

1. Determination of heat capacity of calorimeter for different volumes.
2. Determination of enthalpy of neutralization of hydrochloric acid with sodium hydroxide.
3. Determination of enthalpy of ionization of acetic acid.
4. Determination of integral enthalpy of solution of salts (KNO₃, NH₄Cl).
5. Determination of enthalpy of hydration of copper sulphate.
6. Study of the solubility of benzoic acid in water and determination of ΔH .

Ionic equilibria

pH measurements

- a) Measurement of pH of different solutions like aerated drinks, fruit juices, shampoos and soaps (use dilute solutions of soaps and shampoos to prevent damage to the glass electrode) using pH-meter.
- b) Preparation of buffer solutions:
 - (i) Sodium acetate-acetic acid
 - (ii) Ammonium chloride-ammonium hydroxide

Measurement of the pH of buffer solutions and comparison of the values with theoretical values.

Section B: Organic Chemistry

1. Purification of organic compounds by crystallization (from water and alcohol) and distillation.
2. Criteria of Purity: Determination of melting and boiling points.
3. Preparations: Mechanism of various reactions involved to be discussed. Recrystallisation, determination of melting point and calculation of quantitative yields to be done.
 - (a) Bromination of Phenol/Aniline
 - (b) Benzoylation of amines/phenols
 - (c) Oxime and 2,4-dinitrophenylhydrazone of aldehyde/ketone

Reference Books

- Vogel, A.I., Tatchell, A.R., Furnis, B.S., Hannaford, A.J. & Smith, P.W.G., *Textbook of Practical Organic Chemistry*, Prentice-Hall, 5th edition, 1996.
- Mann, F.G. & Saunders, B.C. *Practical Organic Chemistry* Orient-Longman, 1960.
- Khosla, B. D.; Garg, V. C. & Gulati, A. *Senior Practical Physical Chemistry*, R. Chand & Co.: New Delhi (2011).

DISTRIBUTION OF MARKS

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		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks (Two practicals of 10 marks each)
		Viva-voce: 05 marks

Syllabi and courses of study in chemistry for B.Sc. Semester III for the examinations to be held in Dec. 2017, 2018 & 2019

Course No.: UCHTC 301

Theory: 60 Lectures

**Title: SOLUTIONS, PHASE EQUILIBRIUM, CONDUCTANCE,
ELECTROCHEMISTRY & FUNCTIONAL GROUP ORGANIC
CHEMISTRY**

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Physical Chemistry-II (30 Lectures)

Solutions

(8 Lectures)

Thermodynamics of ideal solutions: Ideal solutions and Raoult's law, deviations from Raoult's law – non-ideal solutions. Vapour pressure-composition and temperature-composition curves of ideal and non-ideal solutions. Azeotropes.

Partial miscibility of liquids: Critical solution temperature; effect of impurity on partial miscibility of liquids. Immiscibility of liquids- Principle of steam distillation. Nernst distribution law and its applications, solvent extraction.

Phase Equilibrium

(8 Lectures)

Phases, components and degrees of freedom of a system, criteria of phase equilibrium. Gibbs Phase Rule and its thermodynamic derivation. Derivation of Clausius – Clapeyron equation and its importance in phase equilibria. Phase diagrams of one-component systems (water and sulphur) and two component systems (lead-silver and $\text{FeCl}_3\text{-H}_2\text{O}$).

Conductance

(6 Lectures)

Conductivity, equivalent and molar conductivity and their variation with dilution for weak and strong electrolytes. Kohlrausch law of independent migration of ions.

Transference number and its experimental determination using Hittorf and Moving boundary methods. Ionic mobility. Applications of conductance measurements: determination of degree of ionization of weak electrolyte, solubility and solubility products of sparingly soluble salts, ionic product of water, hydrolysis constant of a salt. Conductometric titrations (only acid-base).

Electrochemistry

(8 Lectures)

Reversible and irreversible cells. Concept of EMF of a cell. Measurement of EMF of a cell. Nernst equation and its importance. Types of electrodes. Standard electrode potential. Electrochemical series. Thermodynamics of a reversible cell, calculation of thermodynamic properties: G , H , S and equilibrium constant from EMF data.

Concentration cells with transference and without transference. Liquid junction potential and salt bridge.

pH determination using hydrogen electrode and quinhydrone electrode.
Potentiometric titrations -qualitative treatment (acid-base and oxidation-reduction only).

Section B: Organic Chemistry-III (30 Lectures)

Functional group approach for the following reactions (preparations & reactions) to be studied in context to their structure:

Carboxylic acids and their derivatives. (6 Lectures)

Carboxylic acids (aliphatic and aromatic) *Preparation*: Acidic and Alkaline hydrolysis of esters. *Reactions*: Hell-Vohlard-Zelinsky Reaction.

Carboxylic acid derivatives (aliphatic): (Upto 5 carbons)

Preparation: Acid chlorides, Anhydrides, Esters and Amides from acids and their interconversion.

Reactions: Comparative study of nucleophilicity of acyl derivatives. Reformatsky Reaction, Perkin condensation.

Amines and Diazonium Salts (6 Lectures)

Amines (Aliphatic and Aromatic): (Upto 5 carbons)

Preparation: from alkyl halides, Gabriel's Phthalimide synthesis, Hofmann Bromamide reaction.

Reactions: Hofmann vs. Saytzeff elimination, Carbylamine test, Hinsberg test, with HNO_2 , Schotten – Baumann Reaction. Electrophilic substitution (case aniline): nitration, bromination, sulphonation.

Diazonium salts: *Preparation*: from aromatic amines.

Reactions: conversion to benzene, phenol, dyes.

Amino Acids, Peptides and Proteins: (10 Lectures)

Preparation of Amino Acids: Strecker synthesis using Gabriel's phthalimide synthesis. Zwitterion, Isoelectric point and Electrophoresis.

Reactions of Amino acids: ester of $-\text{COOH}$ group, acetylation of $-\text{NH}_2$ group,

complexation with Cu^{2+} ions, ninhydrin test.

Overview of Primary, Secondary, Tertiary and Quaternary Structure of proteins.

Determination of Primary structure of Peptides by degradation Edmann degradation (N-terminal) and C-terminal (thiohydantoin and with carboxypeptidase enzyme). Synthesis of simple peptides (upto dipeptides) by N-protection (*t*-butyloxycarbonyl and phthaloyl) & C-activating groups and Merrifield solid-phase synthesis.

Carbohydrates

(8 Lectures)

Carbohydrates: Classification, and General Properties, Glucose and Fructose (open chain and cyclic structure), Determination of configuration of monosaccharides, absolute configuration of Glucose and Fructose, Mutarotation, ascending and descending in monosaccharides. Structure of disaccharides (sucrose, cellobiose, maltose, lactose) and polysaccharides (starch and cellulose) excluding their structure elucidation.

Reference Books:

- Barrow, G.M. *Physical Chemistry* Tata McGraw-Hill (2007).
- Castellan, G.W. *Physical Chemistry* 4th Ed. Narosa (2004).
- Kotz, J.C., Treichel, P.M. & Townsend, J.R. *General Chemistry*, Cengage Learning India Pvt. Ltd.: New Delhi (2009).
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- Petrucci, R.H. *General Chemistry*, 5th Ed., Macmillan Publishing Co.: New York (1985).
- Morrison, R. T. & Boyd, R. N. *Organic Chemistry*, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
- Finar, I. L. *Organic Chemistry (Volume 1)*, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
- Finar, I. L. *Organic Chemistry (Volume 2)*, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
- Nelson, D. L. & Cox, M. M. *Lehninger's Principles of Biochemistry 7th Ed.*, W. H. Freeman.
- Berg, J.M., Tymoczko, J.L. & Stryer, L. *Biochemistry*, W.H. Freeman, 2002.

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

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SEMESTER-III

Course No.: UCHPC 301 (Practicals)

(For examinations to be held in Dec. 2017, 2018 & 2019)

Title:- Laboratory Course: Chemistry-III

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

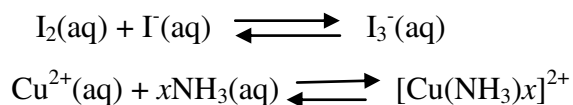
Time: 4 Hrs

Internal Examination: 25 marks

Section A: Physical Chemistry

Distribution

Study of the equilibrium of one of the following reactions by the distribution method:



Phase equilibria

- Construction of the phase diagram of a binary system (simple eutectic) using cooling curves.
- Determination of the critical solution temperature and composition of the phenol water system and study of the effect of impurities on it.
- Study of the variation of mutual solubility temperature with concentration for the phenol water system and determination of the critical solubility temperature.

Conductance

- Determination of cell constant
- Determination of equivalent conductance, degree of dissociation and dissociation constant of a weak acid.
- Perform the following conductometric titrations:
 - Strong acid vs. strong base
 - Weak acid vs. strong base

Potentiometry

Perform the following potentiometric titrations:

- Strong acid vs. strong base
- Weak acid vs. strong base
- Potassium dichromate vs. Mohr's salt

Section B: Organic Chemistry

I Systematic Qualitative Organic Analysis of Organic Compounds possessing monofunctional groups (-COOH, phenolic, aldehydic, ketonic, amide, nitro, amines) and preparation of one derivative.

II

- Determination of the concentration of glycine solution by formylation method.

2. Titration curve of glycine
3. Action of salivary amylase on starch
4. Effect of temperature on the action of salivary amylase on starch.
5. Differentiation between a reducing and a nonreducing sugar.

Reference Books:

- Vogel, A.I., Tatchell, A.R., Furnis, B.S., Hannaford, A.J. & Smith, P.W.G., *Textbook of Practical Organic Chemistry*, Prentice-Hall, 5th edition, 1996.
- Mann, F.G. & Saunders, B.C. *Practical Organic Chemistry* Orient-Longman, 1960.
- Khosla, B. D.; Garg, V. C. & Gulati, A. *Senior Practical Physical Chemistry*, R. Chand & Co.: New Delhi (2011).
- Ahluwalia, V.K. & Aggarwal, R. *Comprehensive Practical Organic Chemistry*, Universities Press.

DISTRIBUTION OF MARKS

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		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks (Two practicals of 10 marks each)
		Viva-voce: 05 marks

SEMESTER-III

(For examinations to be held in Dec. 2017, 2018 & 2019)

Title: FUEL CHEMISTRY (Skill Enhancement Course SEC-1)

Course No.: UCHTS 301

Maximum Marks: 100

Credits: 04

External Examination: 80 marks

Time: 03 Hrs

Internal Assessment: 20 marks

Review of energy sources (renewable and non-renewable). Classification of fuels and their calorific value.

Coal: Uses of coal (fuel and nonfuel) in various industries, its composition, carbonization of coal. Coal gas, producer gas and water gas—composition and uses. Fractionation of coal tar, uses of coal tar based chemicals, requisites of a good metallurgical coke, Coal gasification (Hydro gasification and Catalytic gasification), Coal liquefaction and Solvent Refining.

Petroleum and Petrochemical Industry: Composition of crude petroleum, Refining and different types of petroleum products and their applications.

Fractional Distillation (Principle and process), Cracking (Thermal and catalytic cracking), Reforming Petroleum and non-petroleum fuels (LPG, CNG, LNG, bio-gas, fuels derived from biomass), fuel from waste, synthetic fuels (gaseous and liquids), clean fuels. Petrochemicals: Vinyl acetate, Propylene oxide, Isoprene, Butadiene, Toluene and its derivatives Xylene.

Lubricants: Classification of lubricants, lubricating oils (conducting and non-conducting) Solid and semisolid lubricants, synthetic lubricants.

Properties of lubricants (viscosity index, cloud point, pour point) and their determination.

Reference Books:

- Stocchi, E. *Industrial Chemistry*, Vol-I, Ellis Horwood Ltd. UK (1990).
- Jain, P.C. & Jain, M. *Engineering Chemistry* Dhanpat Rai & Sons, Delhi.
- Sharma, B.K. & Gaur, H. *Industrial Chemistry*, Goel Publishing House, Meerut (1996).

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
- Weightage for attendance – 5 marks

External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

The question paper will have three sections.

Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

Section B will consist of five medium answer type questions (with explanation having 250 to 300 words) of seven marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 35 marks).

Section C will consist of five long answer type questions (with detailed explanation having 500 to 600 words) of fifteen marks each, covering the entire syllabus. The candidate shall have to attempt any two questions (Total marks: 30).

GOVERNMENT COLLEGE FOR WOMEN, PARADE GROUND, JAMMU
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Syllabi and courses of study in chemistry for B.Sc. Semester IV for the examinations to be held in May 2018, 2019 & 2020

Course No.: UCHTC 401

Theory: 60 Lectures

Title: COORDINATION CHEMISTRY, STATES OF MATTER & CHEMICAL KINETICS

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Inorganic Chemistry-II (30 Lectures)

Transition Elements (3d series)

(12 Lectures)

General group trends with special reference to electronic configuration, variable valency, colour, magnetic and catalytic properties, ability to form complexes and stability of various oxidation states (Latimer diagrams) for Mn, Fe and Cu.

Lanthanoids and actinoids: Electronic configurations, oxidation states, colour, magnetic properties, lanthanide contraction, separation of lanthanides (ion exchange method only).

Coordination Chemistry

(8 Lectures)

Valence Bond Theory (VBT): Inner and outer orbital complexes of Cr, Fe, Co, Ni and Cu (coordination numbers 4 and 6). Structural and stereoisomerism in complexes with coordination numbers 4 and 6.

Drawbacks of VBT. IUPAC system of nomenclature.

Crystal Field Theory

(10 Lectures)

Crystal field effect, octahedral symmetry. Crystal field stabilization energy (CFSE), Crystal field effects for weak and strong fields. Tetrahedral symmetry. Factors affecting the magnitude of D. Spectrochemical series. Comparison of CFSE for O_h and T_d complexes, Tetragonal distortion of octahedral geometry.

Jahn-Teller distortion, Square planar coordination.

Section B: Physical Chemistry-III (30 Lectures)

Kinetic Theory of Gases

(8 Lectures)

Postulates of Kinetic Theory of Gases and derivation of the kinetic gas equation.

Deviation of real gases from ideal behaviour, compressibility factor, causes of deviation. van der Waals equation of state for real gases. Boyle temperature (derivation not required). Critical phenomena, critical constants and their calculation from van der Waals equation. Andrews isotherms of CO₂.

Maxwell Boltzmann distribution laws of molecular velocities and molecular energies (graphic representation – derivation not required) and their importance.

Temperature dependence of these distributions. Most probable, average and root mean square velocities (no derivation). Collision cross section, collision number, collision frequency, collision diameter and mean free path of molecules. Viscosity of gases and effect of temperature and pressure on coefficient of viscosity (qualitative treatment only).

Liquids

(6 Lectures)

Surface tension and its determination using stalagmometer. Viscosity of a liquid and determination of coefficient of viscosity using Ostwald viscometer. Effect of temperature on surface tension and coefficient of viscosity of a liquid (qualitative treatment only).

Liquid Crystals: Introduction and types (nematic, smectic & cholesteric)

Solids

(8 Lectures)

Forms of solids. Symmetry elements, unit cells, crystal systems, Bravais lattice types and identification of lattice planes. Laws of Crystallography - Law of constancy of interfacial angles, Law of rational indices. Miller indices. X-Ray diffraction by crystals, Bragg's law. Structures of NaCl, KCl and CsCl (qualitative treatment only). Defects in crystals.

Chemical Kinetics

(8 Lectures)

The concept of reaction rates. Effect of temperature, pressure, catalyst and other factors on reaction rates. Order and molecularity of a reaction. Derivation of integrated rate equations for zero, first and second order reactions (both for equal and unequal concentrations of reactants). Half-life of a reaction. General methods for determination of order of a reaction. Concept of activation energy and its calculation from Arrhenius equation.

Theories of Reaction Rates: Collision theory and Activated Complex theory of bimolecular reactions. Comparison of the two theories (qualitative treatment only).

Reference Books:

- Barrow, G.M. *Physical Chemistry* Tata McGraw-Hill (2007).
- Castellan, G.W. *Physical Chemistry* 4th Ed. Narosa (2004).
- Kotz, J.C., Treichel, P.M. & Townsend, J.R. *General Chemistry* Cengage Learning India Pvt. Ltd., New Delhi (2009).
- Mahan, B.H. *University Chemistry* 3rd Ed. Narosa (1998).
- Petrucci, R.H. *General Chemistry* 5th Ed. Macmillan Publishing Co.: New York (1985).
- Cotton, F.A. & Wilkinson, G. *Basic Inorganic Chemistry*, Wiley.
- Shriver, D.F. & Atkins, P.W. *Inorganic Chemistry*, Oxford University Press.
- Wulfsberg, G. *Inorganic Chemistry*, Viva Books Pvt. Ltd.
- Rodgers, G.E. *Inorganic & Solid State Chemistry*, Cengage Learning India Ltd., 2008.

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
- Weightage for attendance – 5 marks

External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

The question paper will have three sections.

Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

Section B will consist of five medium answer type questions (with explanation having 250 to 300 words) of seven marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 35 marks).

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SEMESTER-IV

Course No.: UCHPC 401 (Practicals)

(For examinations to be held in May 2018, 2019 & 2020)

Title:- Laboratory Course: Chemistry-I

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Inorganic Chemistry

Semi-micro qualitative analysis using H₂S of mixtures - not more than four ionic species (two anions and two cations and excluding insoluble salts) out of the following:

Cations : NH₄⁺, Pb²⁺, Ag⁺, Bi³⁺, Cu²⁺, Cd²⁺, Sn²⁺, Fe³⁺, Al³⁺, Co²⁺, Cr³⁺, Ni²⁺, Mn²⁺, Zn²⁺, Ba²⁺, Sr²⁺, Ca²⁺, K⁺

Anions : CO₃²⁻, S²⁻, SO₃²⁻, S₂O₃²⁻, NO₂⁻, CH₃COO⁻, Cl⁻, Br⁻, I⁻, NO₃⁻, SO₄²⁻, PO₄³⁻, BO₃³⁻, C₂O₄²⁻, F⁻

(Spot tests should be carried out wherever feasible)

1. Estimate the amount of nickel present in a given solution as bis(dimethylglyoximato) nickel(II) or aluminium as oximate in a given solution gravimetrically.
2. Draw calibration curve (absorbance at λ_{\max} vs. concentration) for various concentrations of a given coloured compound (KMnO₄/ CuSO₄) and estimate the concentration of the same in a given solution.
3. Determine the composition of the Fe³⁺-salicylic acid complex solution by Job's method.
4. Estimation of (i) Mg²⁺ or (ii) Zn²⁺ by complexometric titrations using EDTA.
5. Estimation of total hardness of a given sample of water by complexometric titration.
6. Determination of concentration of Na⁺ and K⁺ using Flame Photometry.

Section B: Physical Chemistry

(I) Surface tension measurement (use of organic solvents excluded).

- a) Determination of the surface tension of a liquid or a dilute solution using a stalagmometer.
- b) Study of the variation of surface tension of a detergent solution with concentration.

(II) Viscosity measurement (use of organic solvents excluded).

- a) Determination of the relative and absolute viscosity of a liquid or dilute solution using an Ostwald's viscometer.
- b) Study of the variation of viscosity of an aqueous solution with concentration of solute.

(III) Chemical Kinetics

Study the kinetics of the following reactions.

1. Initial rate method: Iodide-persulphate reaction
2. Integrated rate method:
 - a. Acid hydrolysis of methyl acetate with hydrochloric acid.
 - b. Saponification of ethyl acetate.
 - c. Compare the strengths of HCl and H₂SO₄ by studying kinetics of hydrolysis of methyl acetate

Reference Books:

- Svehla, G. *Vogel's Qualitative Inorganic Analysis*, Pearson Education, 2012.
- Mendham, J. *Vogel's Quantitative Chemical Analysis*, Pearson, 2009.
- Khosla, B. D.; Garg, V. C. & Gulati, A. *Senior Practical Physical Chemistry*, R. Chand & Co.: New Delhi (2011).

DISTRIBUTION OF MARKS

A.	Daily evaluation of practical records/viva-voce/attendance, etc.	Attendance: 05 marks
		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks
		(Two practicals of 10 marks each)
		Viva-voce: 05 marks

SEMESTER-IV

(For examinations to be held in May 2018, 2019 & 2020)

Title: PESTICIDE CHEMISTRY (Skill Enhancement Course SEC-2)

Course No.: UCHTS 401

Maximum Marks: 100

Credits: 04

External Examination: 80 marks

Time: 03 Hrs

Internal Assessment: 20 marks

General introduction to pesticides (natural and synthetic), benefits and adverse effects, changing concepts of pesticides, structure activity relationship, synthesis and technical manufacture and uses of representative pesticides in the following classes: Organochlorines (DDT, Gammexene); Organophosphates (Malathion, Parathion); Carbamates (Carbofuran and carbaryl); Quinones (Chloranil), Anilides (Alachlor and Butachlor).

Practicals

- To calculate acidity/alkalinity in given sample of pesticide formulations as per BIS specifications.
- Preparation of simple organophosphates, phosphonates and thiophosphates

Reference Book:

1. Cremllyn, R. *Pesticides. Preparation and Modes of Action*, John Wiley & Sons, New York, 1978.

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
- Weightage for attendance – 5 marks

External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

The question paper will have three sections.

Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

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Syllabi and courses of study in chemistry for B.Sc. Semester V for the examinations to be held in Dec. 2018, 2019 & 2020

Course No. :- UCHTDSE 501

Theory: 60 Lectures

Title:- SPECTROSCOPY, PHOTOCHEMISTRY, ORGANOMETALLICS AND BIOINORGANIC CHEMISTRY

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Physical Chemistry-IV

Molecular Spectroscopy: (10 Lectures)

Interaction of electromagnetic radiation with molecules and various types of spectra; Born-Oppenheimer approximation.

Rotational spectroscopy: Selection rules, intensities of spectral lines, determination of bond lengths of diatomic and linear triatomic molecules, isotopic substitution.

Vibrational spectroscopy: Classical equation of vibration, computation of force constant, amplitude of diatomic molecular vibrations, anharmonicity, Morse potential, dissociation energies, fundamental frequencies, overtones, hot bands, degrees of freedom for polyatomic molecules.

Photochemistry (10 Lectures)

Characteristics of electromagnetic radiation, Lambert-Beer's law and its limitations, physical significance of absorption coefficients. Laws of photochemistry, quantum yield, photophysical process (Jablonski diagram), photosensitised reactions, quenching.

Physical properties and Molecular Structure: (10 Lectures)

Optical activity, Polarization – (Clausius-Mossotti equation), Orientation of dipoles in an electric field, Dipole moment, Induced dipole moment, Measurement of dipole moment: temperature method and refractivity method, Dipole moment and structure of molecules, Magnetic properties: paramagnetism, diamagnetism and ferromagnetism.

Section B: Inorganic Chemistry-III

Chemistry of 3d metals

(6 Lectures)

Oxidation states displayed by Cr, Fe, Co, Ni and Cu.

A study of the following compounds (including preparation and important properties);

Peroxo compounds of Cr, $K_2Cr_2O_7$, $KMnO_4$, $K_4[Fe(CN)_6]$, sodium nitroprusside, $[Co(NH_3)_6]Cl_3$, $Na_3[Co(NO_2)_6]$.

Organometallic Compounds

(12 Lectures)

Definition and classification with appropriate examples based on nature of metal-carbon bond (ionic, σ , π , and multicentre bonds). Structures of methyl lithium, Zeise's salt and ferrocene. EAN rule as applied to carbonyls. Preparation, structure, bonding and properties of mononuclear and polynuclear carbonyls of 3d metals. π -acceptor behaviour of carbon monoxide. Synergic effects (VB approach)- (MO diagram of CO can be referred to for synergic effect to IR frequencies).

Bio-Inorganic Chemistry

(12 Lectures)

A brief introduction to bio-inorganic chemistry. Role of metal ions present in biological systems with special reference to Na^+ , K^+ and Mg^{2+} ions: Na/K pump; Role of Mg^{2+} ions in energy production and chlorophyll. Role of Ca^{2+} in blood clotting, stabilization of protein structures and structural role (bones).

Reference Books:

- 1) Banwell, C. N. & McCash, E. M. *Fundamentals of Molecular Spectroscopy* 4th Ed. Tata McGraw-Hill: New Delhi (2006).
- 2) Kakkar, R. *Atomic & Molecular Spectroscopy: Concepts & Applications*, Cambridge University Press (2015).
- 3) James E. Huheey, Ellen Keiter & Richard Keiter: *Inorganic Chemistry: Principles of Structure and Reactivity*, Pearson Publication.
- 4) G.L. Miessler & Donald A. Tarr: *Inorganic Chemistry*, Pearson Publication.
- 5) J.D. Lee: *A New Concise Inorganic Chemistry*, E.L.B.S.
- 6) F.A. Cotton & G. Wilkinson: *Basic Inorganic Chemistry*, John Wiley & Sons

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
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External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

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Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

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SEMESTER-V

Course No. :- UCHPDSE 501 (Practicals)

(For examinations to be held in Dec. 2018, 2019 & 2020)

Title:- Laboratory Course: Chemistry-V

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Physical Chemistry - Colourimetry

- I. Verify Lambert-Beer's law and determine the concentration of $\text{CuSO}_4/\text{KMnO}_4/\text{K}_2\text{Cr}_2\text{O}_7$ in a solution of unknown concentration.
- II. Determine the concentrations of KMnO_4 and $\text{K}_2\text{Cr}_2\text{O}_7$ in a mixture.
- III. Determine the dissociation constant of an indicator (phenolphthalein).
- IV. Analyse the given vibration-rotation spectrum of $\text{HCl}(\text{g})$.

Section B: Inorganic Chemistry - UV/Visible spectroscopy

- 1) Study the 200-500 nm absorbance spectra of KMnO_4 and $\text{K}_2\text{Cr}_2\text{O}_7$ (in 0.1 M H_2SO_4) and determine the λ_{max} values. Calculate the energies of the two transitions in different units (J molecule^{-1} , kJ mol^{-1} , cm^{-1} , eV).
- 2) Study the pH-dependence of the UV-Vis spectrum (200-500 nm) of $\text{K}_2\text{Cr}_2\text{O}_7$.
- 3) Record the 200-350 nm UV spectra of the given compounds (acetone, acetaldehyde, 2-propanol, acetic acid) in water. Comment on the effect of structure on the UV spectra of organic compounds.

Reference Books

- 1) Mendham, J. *Vogel's Quantitative Chemical Analysis*, Pearson, 2009.
- 2) Khosla, B. D.; Garg, V. C. & Gulati, A., *Senior Practical Physical Chemistry*, R. Chand & Co.: New Delhi (2011).
- 3) Garland, C. W.; Nibler, J. W. & Shoemaker, D. P. *Experiments in Physical Chemistry 8th Ed.*; McGraw-Hill: New York (2003).
- 4) Halpern, A. M. & McBane, G. C. *Experimental Physical Chemistry 3rd Ed.*; W.H. Freeman & Co.: New York (2003).
- 5) J.B. Yadav, *Advance Practical Physical Chemistry*, Goel Publication, Meerut.
- 6) B. Vishwanathan and P.S. Raghvan, *Practical Physical Chemistry*, Viva Books Pvt. Ltd.

DISTRIBUTION OF MARKS

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		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks
		(Two practicals of 10 marks each)
		Viva-voce: 05 marks

GOVERNMENT COLLEGE FOR WOMEN, PARADE GROUND, JAMMU
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Syllabi and courses of study in chemistry for B.Sc. Semester V for the examinations to be held in Dec. 2018, 2019 & 2020

Course No.: UCHTDSE 511

Theory: 60 Lectures

Title: QUANTUM MECHANICS, POLYMERIC MATERIALS, INDUSTRIAL METALLURGY & ENVIRONMENT

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Physical Chemistry-IV (30 lectures)

Quantum Mechanics-I

(18 lectures)

Black-body radiation and Planck's radiation law, Photoelectric effect, Heat capacity of solids, Bohr's model of hydrogen atom and its defects, Compton effect, Zeeman effect, de Broglie hypothesis, Heisenberg's uncertainty principle.

Quantum Mechanics-II

Schrodinger wave equation and its importance, Physical interpretation of the wave function, Concept of operators, Hamiltonian operator, Postulates of quantum mechanics, Discussion of solutions of the Schrodinger equation to some model systems viz; Particle in one dimensional box, Particle in three dimensional box, Concept of degeneracy.

Introduction and history of polymeric materials:

(12 lectures)

Different schemes of classification of polymers, Polymer nomenclature, Molecular forces and chemical bonding in polymers,

Kinetics of Polymerization:

Mechanism and kinetics of step growth, radical chain growth, ionic chain (both cationic and anionic) and coordination polymerizations, Mechanism and kinetics of copolymerization, polymerization techniques

Determination of molecular weight of polymers (M_n , M_w , etc) by end group analysis, viscometry, light scattering and osmotic pressure methods. Molecular weight distribution and its significance. Polydispersity index.

Section B: Inorganic Chemistry-III (30 lectures)

Industrial Metallurgy - General Principles of Metallurgy

(10 lectures)

Chief modes of occurrence of metals based on standard electrode potentials. Ellingham diagrams for reduction of metal oxides using carbon as reducing agent.

Hydrometallurgy, Methods of purification of metals (Al, Pb, Ti, Fe, Cu, Ni, Zn): electrolytic, oxidative refining, Kroll process, Parting process, van Arkel-de Boer process and Mond's process.

Preparation of metals (ferrous and nonferrous) and ultrapure metals for semiconductor technology.

Environment and its segments

(20 lectures)

Ecosystems. Biogeochemical cycles of carbon, nitrogen and sulphur.

Air Pollution: Major regions of atmosphere. Chemical and photochemical reactions in atmosphere. Air pollutants: types, sources, particle size and chemical nature; Photochemical smog: its constituents and photochemistry. Environmental effects of ozone, Major sources of air pollution. Pollution by SO₂, CO₂, CO, NO_x, H₂S and other foul smelling gases. Methods of estimation of CO, NO_x, SO_x and control procedures. Effects of air pollution on living organisms and vegetation. Greenhouse effect and Global warming, Ozone depletion by oxides of nitrogen, chlorofluorocarbons and Halogens, removal of sulphur from coal. Control of particulates.

Water Pollution : Hydrological cycle, water resources, aquatic ecosystems, Sources and nature of water pollutants, Techniques for measuring water pollution, Impact of water pollution on hydrological and ecosystems.

Water purification methods. Effluent treatment plants (primary, secondary and tertiary treatment). Industrial effluents from the following industries and their treatment: electroplating, textile, tannery, dairy, petroleum and petrochemicals, agro, fertilizer, etc.

Reference Books:

- 1) Chandra, A. K. *Introductory Quantum Chemistry* Tata McGraw-Hill (2001).
- 2) House, J. E. *Fundamentals of Quantum Chemistry* 2nd Ed. Elsevier: USA (2004).
- 3) Lowe, J. P. & Peterson, K. *Quantum Chemistry*, Academic Press (2005).
- 4) Seymour, R.B. & Carraher, C.E. *Polymer Chemistry: An Introduction*, Marcel Dekker, Inc. New York, 1981.
- 5) Odian, G. *Principles of Polymerization*, 4th Ed. Wiley, 2004.
- 6) Billmeyer, F.W. *Textbook of Polymer Science*, 2nd Ed. Wiley Interscience, 1971.
- 7) Ghosh, P. *Polymer Science & Technology*, Tata McGraw-Hill Education, 1991
- 8) P. Munk & T.M. Aminabhavi, *Introduction to Macromolecular Science*, 2nd ed. John Wiley & Sons (2002)
- 9) E. Stocchi: *Industrial Chemistry*, Vol-I, Ellis Horwood Ltd. UK.
- 10) R.M. Felder, R.W. Rousseau: *Elementary Principles of Chemical Processes*, Wiley Publishers, New Delhi.

- 11) J. A. Kent: *Riegel's Handbook of Industrial Chemistry*, CBS Publishers, New Delhi.
 - 12) S. S. Dara: *A Textbook of Engineering Chemistry*, S. Chand & Company Ltd. New Delhi.
 - 13) K. De, *Environmental Chemistry*: New Age International Pvt., Ltd, New Delhi.
 - 14) S. M. Khopkar, *Environmental Pollution Analysis*: Wiley Eastern Ltd, New Delhi.
 - 15) S.E. Manahan, *Environmental Chemistry*, CRC Press (2005).
 - 16) G.T. Miller, *Environmental Science* 11th edition. Brooks/ Cole (2006).
 - 17) A. Mishra, *Environmental Studies*. Selective and Scientific Books, New Delhi (2005).
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NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

- The internal assessment test shall be held on completion of about 40% of the prescribed syllabus. The question paper will consist of eight short answer type questions of three marks each. The candidate will have to attempt any five questions (Total: 15 marks).
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External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

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SEMESTER-V

Course No. :- UCHPDSE 511 (Practicals)

(For examinations to be held in Dec. 2018, 2019 & 2020)

Title:- Laboratory Course: Chemistry-V

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Physical Chemistry

- 1) Determination of molecular weight of different polymers by viscometry:
 - (a) Polyacrylamide-aq. NaNO₂ solution
 - (b) Poly vinyl propylidene (PVP) in water
- 2) Determination of the viscosity-average molecular weight of poly(vinyl alcohol) (PVOH) and the fraction of “head-to-head” monomer linkages in the polymer.
- 3) Estimation of the amount of HCHO in the given solution by sodium sulphite method
- 4) Testing of mechanical properties of polymers.
- 5) Determination of molecular weight by end group analysis: Polyethylene glycol (PEG) (OH group).
- 6) Determination of hydroxyl number of a polymer using colorimetric method.

Section B: Inorganic Chemistry

- 1) Determination of dissolved oxygen in water.
- 2) Determination of Chemical Oxygen Demand (COD)
- 3) Determination of Biological Oxygen Demand (BOD)
- 4) Percentage of available chlorine in bleaching powder.
- 5) Measurement of chloride, sulphate and salinity of water samples by simple titration method (AgNO₃ and potassium chromate).
- 6) Estimation of total alkalinity of water samples (CO₃²⁻, HCO₃⁻) using double titration method.

Reference Books

- 1) M.P. Stevens, *Polymer Chemistry: An Introduction*, 3rd Ed., Oxford University Press, 1999.
- 2) H.R. Allcock, F.W. Lampe & J.E. Mark, *Contemporary Polymer Chemistry*, 3rd ed. Prentice-Hall (2003)
- 3) F.W. Billmeyer, *Textbook of Polymer Science*, 3rd ed. Wiley-Interscience (1984)
- 4) J.R. Fried, *Polymer Science and Technology*, 2nd ed. Prentice-Hall (2003)
- 5) P. Munk & T.M. Aminabhavi, *Introduction to Macromolecular Science*, 2nd ed. John Wiley & Sons (2002)
- 6) L. H. Sperling, *Introduction to Physical Polymer Science*, 4th ed. John Wiley & Sons (2005)
- 7) M.P. Stevens, *Polymer Chemistry: An Introduction* 3rd ed. Oxford University Press (2005).
- 8) Seymour/ Carraher, *Polymer Chemistry*, 9th ed. by Charles E. Carraher, Jr. (2013).
- 9) E. Stocchi: *Industrial Chemistry*, Vol-I, Ellis Horwood Ltd. UK.
- 10) R.M. Felder, R.W. Rousseau: *Elementary Principles of Chemical Processes*, Wiley Publishers, New Delhi.
- 11) J. A. Kent: *Riegel's Handbook of Industrial Chemistry*, CBS Publishers, New Delhi.
- 12) S. S. Dara: *A Textbook of Engineering Chemistry*, S. Chand & Company Ltd. New Delhi.
- 13) A. K. De, *Environmental Chemistry*: New Age International Pvt. Ltd, New Delhi.

14) S. M. Khopkar, *Environmental Pollution Analysis*: Wiley Eastern Ltd, New Delhi.

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A.	Daily evaluation of practical records/viva-voce/attendance, etc.	Attendance: 05 marks
		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks (Two practicals of 10 marks each)
		Viva-voce: 05 marks

SEMESTER-V

(For examinations to be held in Dec. 2018, 2019 & 2020)

Title: COSMETICS, PERFUMES AND MEDICINAL AGENTS FROM NATURAL SOURCES (Skill Enhancement Course SEC-3)

Course No.: UCHTS 501

Maximum Marks: 100

Credits: 04

External Examination: 80 marks

Time: 03 Hrs

Internal Assessment: 20 marks

A general study including preparation and uses of the following: Hair dye, hair spray, shampoo, suntan lotions, face powder, lipsticks, talcum powder, nail enamel, creams (cold, vanishing and shaving creams), antiperspirants and artificial flavours. Essential oils and their importance in cosmetic industries with reference to Eugenol, Geraniol, sandalwood oil, eucalyptus, rose oil, 2-phenyl ethyl alcohol, Jasmine, Civetone, Muscone.

Different systems of classification of drugs of natural origin, their merits & demerits.

Study of mineral ingredients. Kaolin, Bentonite, Talc., Fuller's earth.

Study of the following with reference to their sources, characters, chemical constituents, identification tests and cosmetic uses.

i) Carbohydrate – Starches, Agar, Gum Acacia.

ii) Lipids – A study of following; Caster oil, Linseed oil, Sesame oil, Coconut oil.

iii) Wheat germ. Oil, Jojoba oil, rice bran oil, Spermaceti.

Cosmeceuticals – study of hair care herbs and skin care herbs as active ingredients.

Tannins – Classification, identification and study of the following – Black Catechu, Tannic Acid, Amla, Behra, Hirida, Arjun, Ashok.

Practicals

1. Preparation of talcum powder.
2. Preparation of shampoo.
3. Preparation of enamels.
4. Preparation of hair remover.
5. Preparation of face cream.
6. Preparation of nail polish and nail polish remover.

Reference Books:

- 1) E. Stocchi: Industrial Chemistry, Vol -I, Ellis Horwood Ltd. UK.
- 2) P.C. Jain, M. Jain: Engineering Chemistry, Dhanpat Rai & Sons, Delhi.
- 3) Sharma, B.K. & Gaur, H. Industrial Chemistry, Goel Publishing House, Meerut(1996).
- 4) Text book of Pharmacognosy – Trease and Evan's.
- 5) Pharmacognosy – By Clasu and Tayler.
- 6) Text Book of Pharmacognosy – T.E. Wallis.
- 7) Materia Medica – By Nadkarni.

NOTE FOR PAPER SETTERS

Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

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- Weightage for attendance – 5 marks

External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

The question paper will have three sections.

Section A will consist of five short answer type questions (with explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

Section B will consist of five medium answer type questions (with explanation having 250 to 300 words) of seven marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 35 marks).

Section C will consist of five long answer type questions (with detailed explanation having 500 to 600 words) of fifteen marks each, covering the entire syllabus. The candidate shall have to attempt any two questions (Total marks: 30).

Syllabi and courses of study in chemistry for B.Sc. Semester VI for the examinations to be held in May 2019, 2020 & 2021

Course No.: UCHTDSE 601

Theory: 60 Lectures

Title:- INORGANIC MATERIALS OF INDUSTRIAL IMPORTANCE AND ORGANIC SPECTROSCOPY

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Inorganic Chemistry-IV (30 Lectures)

Recapitulation of *s*- and *p*-Block Elements

(4 Lectures)

Periodicity in *s*- and *p*-block elements with respect to electronic configuration, atomic and ionic size, ionization enthalpy, electronegativity (Pauling, Mulliken, and Alfred-Rochow scales). Allotropy in C, S, and P.

Oxidation states with reference to elements in unusual and rare oxidation states like carbides and nitrides), inert pair effect, diagonal relationship and anomalous behaviour of first member of each group.

Fertilizers:

(8 Lectures)

Different types of fertilizers. Manufacture of the following fertilizers: Urea, ammonium nitrate, calcium ammonium nitrate, ammonium phosphates; polyphosphate, superphosphate, compound and mixed fertilizers, potassium chloride, potassium sulphate.

Batteries:

(8 Lectures)

Primary and secondary batteries, battery components and their role, Characteristics of Battery. Working of following batteries: Pb acid battery, Li-battery, Solid state electrolyte battery. Fuel cells, Solar cell and polymer cell.

Catalysis:

(6 Lectures)

General principles and properties of catalysts, homogenous catalysis (catalytic steps and examples) and heterogenous catalysis (catalytic steps and examples) and their industrial applications, Deactivation or regeneration of catalysts.

Chemical explosives:

(4 Lectures)

Origin of explosive properties in organic compounds, preparation and explosive properties of lead azide, PETN, cyclonite (RDX). Introduction to rocket propellants.

Section B: Organic Chemistry-IV (30 Lectures)

Active methylene compounds: (8 Lectures)

Preparation: Claisen ester condensation. Keto-enol tautomerism.

Reactions: Synthetic uses of ethylacetoacetate (2-hexanone, Pentanoic acid and β -ketoester).

Application of Spectroscopy to Simple Organic Molecules

a) Electromagnetic spectrum (12 Lectures)

Ultraviolet (UV) absorption spectroscopy – absorption laws (Beer Lambert law), molar absorptivity, types of electronic transitions, effect of conjugation. Concept of Chromophore and auxochrome. Bathochromic, hypsochromic, hyperchromic and hypochromic shifts, effect of solvents. UV spectra of conjugated enes and enones (Woodward - Fieser rules).

Infrared (IR) absorption spectroscopy – molecular vibrations. Hooke's law, selection rules, intensity and position of IR bands, fingerprint region, characteristic absorptions of various functional groups and interpretation of IR spectra of simple organic compounds (benzaldehyde, benzoyl chloride, acetophenone, benzamide, salicylic acid, p-hydroxybenzoic acid, p-nitroaniline, p-nitrophenol, benzonitrile, α,β -unsaturated ketones).

b) Spectroscopy (10 Lectures)

Nuclear magnetic resonance (NMR) spectroscopy. Proton magnetic resonance (^1H NMR) spectroscopy, nuclear shielding and deshielding, chemical shift and molecular structure, spin-spin splitting and coupling constants, areas of signals, interpretation of PMR spectra of simple organic molecules such as ethyl bromide, ethanol, acetaldehyde, 1,1,2-tribromoethane, ethyl acetate, toluene and acetophenone.

Reference Books:

- 1) R. M. Felder, R. W. Rousseau: *Elementary Principles of Chemical Processes*, Wiley Publishers, New Delhi.
- 2) J. A. Kent: *Riegel's Handbook of Industrial Chemistry*, CBS Publishers, New Delhi.
- 3) R. Gopalan, D. Venkappayya, S. Nagarajan: *Engineering Chemistry*, Vikas Publications, New Delhi.
- 4) B. K. Sharma: *Engineering Chemistry*, Goel Publishing House, Meerut.
- 5) J.D. Lee: *A New Concise Inorganic Chemistry*, E.L.B.S.
- 6) F.A. Cotton & G. Wilkinson: *Basic Inorganic Chemistry*, John Wiley & Sons.
- 7) I.L. Finar: *Organic Chemistry* (Vol. I & II), E.L.B.S.

- 8) John R. Dyer: *Applications of Absorption Spectroscopy of Organic Compounds*, Prentice Hall.
- 9) R.M. Silverstein, G.C. Bassler & T.C. Morrill: *Spectroscopic Identification of Organic Compounds*, John Wiley & Sons.
- 10) R.T. Morrison & R.N. Boyd: *Organic Chemistry*, Prentice Hall.
- 11) Peter Sykes: *A Guide Book to Mechanism in Organic Chemistry*, Orient Longman.
- 12) Arun Bahl and B. S. Bahl: *Advanced Organic Chemistry*, S. Chand.
- 13) V.K. Ahluwalia and R.K. Parashar: *Organic Reaction Mechanism*, Narosa.

NOTE FOR PAPER SETTERS

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External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

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Section A will consist of five short answers type questions (without detailed explanation having 70 to 80 words) of three marks each, covering the entire syllabus. The candidate shall have to attempt all the questions (Total: 15 marks).

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(For examinations to be held in May 2019, 2020 & 2021)

Course No.: UCHPDSE 601 (Practicals)

Title:- Laboratory Course: Chemistry-VI

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Inorganic Chemistry

- 1) Determination of free acidity in ammonium sulphate fertilizer.
- 2) Estimation of calcium in calcium ammonium nitrate fertilizer.
- 3) Estimation of phosphoric acid in superphosphate fertilizer.
- 4) Electroless metallic coatings on ceramic and plastic material.
- 5) Determination of composition of dolomite (by complexometric titration).

Section B: Organic Chemistry

- 1) Systematic qualitative analysis of organic compounds and preparation of suitable derivatives (including osazone formation in carbohydrates).
- 2) Separation of binary mixture of organic compounds using water, sodium bicarbonate, sodium hydroxide & hydrochloric acid.

Preparations

- 1) p-Nitroaniline from acetanilide
- 2) p-Bromoaniline from acetanilide
- 3) m-Nitroaniline from nitrobenzene
- 4) Synthesis of di-benzalacetone from benzaldehyde (Claisen-Schmidt condensation)
- 5) Synthesis of benzyl alcohol and benzoic acid from benzaldehyde (Cannizzaro reaction)

Reference Books:

- 1) E. Stocchi: *Industrial Chemistry*, Vol-I, Ellis Horwood Ltd. UK.
- 2) R. M. Felder, R. W. Rousseau: *Elementary Principles of Chemical Processes*, Wiley Publishers, New Delhi.
- 3) W. D. Kingery, H. K. Bowen, D. R. Uhlmann: *Introduction to Ceramics*, Wiley Publishers, New Delhi.
- 4) J. A. Kent: *Riegel's Handbook of Industrial Chemistry*, CBS Publishers, New Delhi.
- 5) P. C. Jain, M. Jain: *Engineering Chemistry*, Dhanpat Rai & Sons, Delhi.
- 6) R. Gopalan, D. Venkappayya, S. Nagarajan: *Engineering Chemistry*, Vikas Publications, New Delhi.
- 7) Mann, F.G. & Saunders, B.C. *Practical Organic Chemistry* Orient-Longman, 1960.
- 8) *Experimental Organic Chemistry – Principles and practice* by Laurence M. Harwood and Christopher J. Moody (Blackwell Scientific Publications).
- 9) *Laboratory Manual of Organic Chemistry* by Raj K. Bansal (New Age Publications).

- 10) Advanced Practical Chemistry by Jagdamba Singh and others (Pragati Prakashan).
- 11) Practical Organic Chemistry by N.K. Vishnoi (New Age Publications).
- 12) Experimental Organic Chemistry, Vol. I & II, P.R. Singh, D.S. Gupta and K.S. Bajpai, Tata McGraw Hill.
- 13) Laboratory Manual in Organic Chemistry, R.K. Bansal, Wiley Eastern.
- 14) Vogel's Textbook of Practical Organic Chemistry, B.S. Furniss, A.J. Hannaford, V. Rogers, P.W.G. Smith and A.R. Tatchell, ELBS.
- 15) Experiments in General Chemistry, C.N.R. Rao and U.C. Agarwal, East-West Press.

DISTRIBUTION OF MARKS

A.	Daily evaluation of practical records/viva-voce/attendance, etc.	Attendance: 05 marks
		Practical Test: 08 marks
		Day to day performance: 12 marks
B.	External Examination (100% syllabus)	Examination: 20 marks
		(Two practicals of 10 marks each)
		Viva-voce: 05 marks

Syllabi and courses of study in chemistry for B.Sc. Semester VI for the examinations to be held in May 2019, 2020 & 2021

Course No.: UCHTDSE 611

Theory: 60 Lectures

Title: ANALYTICAL METHODS, ELECTRONIC AND MAGNETIC PROPERTIES OF TRANSITION METALS AND SYNTHETIC ORGANIC CHEMISTRY

Credits: 04

Maximum Marks: 100

Time: 03 Hrs

External Examination: 80 marks

Internal Assessment: 20 marks

Section A: Inorganic Chemistry-IV (30 Lectures)

Separation techniques: (15 lectures)

Solvent extraction: Classification, principle and efficiency of the technique. Mechanism of extraction: extraction by solvation and chelation. Technique of extraction: batch, continuous and counter current extractions.

Qualitative and quantitative aspects of solvent extraction: extraction of metal ions from aqueous solution, extraction of organic species from the aqueous and nonaqueous media.

Chromatography: Classification, principle and efficiency of the technique. Mechanism of separation: adsorption, partition & ion exchange. Development of chromatograms: frontal, elution and displacement methods. Qualitative and quantitative aspects of chromatographic methods of analysis: IC, GLC, GPC, TLC and HPLC.

Magnetic Properties of Transition Metal Complexes (8 lectures)

Types of Magnetic behaviour, methods of determining magnetic susceptibility(Guoy's and Faraday's methods) spin only formula, L-S coupling and correlation of values, orbital contribution to magnetic moments, application of magnetic moment data for structure analysis of 3d-metal complexes.

Electronic Spectra of Transition Metal Complexes (7 lectures)

Types of electronic transition, selection rules for d- d transitions, spectroscopic ground states, spectro-chemical series, Orgel-energy level diagram for d^1 and d^9 states in octahedral and tetrahedral systems, discussion of the electronic spectrum of $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ complex ion.

Section B: Organic Chemistry-IV (30 Lectures)

Polynuclear and Heteronuclear aromatic compounds:**(10 Lectures)**

Method of synthesis and electrophilic and nucleophilic reactions of naphthalene and anthracene. Molecular orbital picture and aromatic characteristics of pyrrole, furan, thiophene and pyridine. Methods of synthesis and mechanism of electrophilic substitution reactions. Mechanism of nucleophilic substitution reactions in pyridine. Comparison of basicity of pyridine, piperidine and pyrrole.

Organic Synthesis via enolates:**(8 Lectures)**

Acidity of α -hydrogens, alkylation of ethyl acetoacetate. Synthesis of ethyl acetoacetate: the Claisen condensation. Keto-enol tautomerism of ethyl acetoacetate. Alkylation of 1,3-dithianes, alkylation and acylation of enamines.

Organometallic compounds of Mg, Li & Zn**(9 Lectures)**

Organomagnesium compounds: Grignard reagents – formation, structure and chemical reactions. Organozinc compounds: formation and chemical reactions. Organolithium compounds: formation and chemical reactions.

Synthetic dyes**(3 Lectures)**

Synthesis of Methyl orange, Malachite green. Crystal violet, Phenolphthalein and Indigo.

Reference Books:

- 1) Jeffery, G.H., Bassett, J., Mendham, J. & Denney, R.C. *Vogel's Textbook of Quantitative Chemical Analysis*, John Wiley & Sons, 1989.
- 2) Willard, H.H., Merritt, L.L., Dean, J. & Settoe, F.A. *Instrumental Methods of Analysis*, 7th Ed. Wadsworth Publishing Company Ltd., Belmont, California, USA, 1988.
- 3) Christian, G.D; *Analytical Chemistry*, 6th Ed. John Wiley & Sons, New York, 2004.
- 4) Harris, D. C. *Exploring Chemical Analysis*, Ed. New York, W.H. Freeman, 2001.
- 5) Khopkar, S.M. *Basic Concepts of Analytical Chemistry*. New Age, International Publisher, 2009.
- 6) Skoog, D.A. Holler F.J. & Nieman, T.A. *Principles of Instrumental Analysis*, Cengage Learning India Ed.
- 7) Mikes, O. *Laboratory Hand Book of Chromatographic & Allied Methods*, Elles Harwood
- 8) Series on Analytical Chemistry, John Wiley & Sons, 1979.
- 9) Ditts, R.V. *Analytical Chemistry; Methods of Separation*, van Nostrand, 1974.
- 10) I.L. Finar: *Organic Chemistry* (Vol. I & II), E.L.B.S. 52
- 11) R.T. Morrison & R.N. Boyd: *Organic Chemistry*, Prentice Hall.
- 12) Peter Sykes: *A Guide Book to Mechanism in Organic Chemistry*, Orient Longman.
- 13) Arun Bahl and B. S. Bahl: *Advanced Organic Chemistry*, S. Chand.
- 14) Solomons, *Fundamentals of Organic Chemistry*, John Wiley.
- 15) F.A. Carey, *Organic Chemistry*, McGraw Hill Inc.

16) Stritwieser, Healthcock and Kosover, *Introduction to Organic Chemistry*, Macmillan

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Internal Assessment (Total Marks: 20; Time Duration: 1 hour)

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External End Semester Examination (Total Marks: 80; Time Duration: 3 hours)

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SEMESTER-VI

(For examinations to be held in May 2019, 2020 & 2021)

Course No.: UCHPDSE 611 (Practicals)

Title:- Laboratory Course: Chemistry-VI

Maximum Marks: 50

Credits: 02

External Examination: 25 marks

Time: 4 Hrs

Internal Examination: 25 marks

Section A: Inorganic Chemistry

Separation Techniques

1. Chromatography (Separation of Mixture): Paper chromatographic separation of Fe^{3+} , Al^{3+} , and Cr^{3+} .
- 2) To separate a mixture of Ni^{2+} & Fe^{2+} by complexation with DMG and extracting the Ni^{2+} DMG complex in chloroform, and determine its concentration by spectrophotometry.
- 3) Analysis of soil.
 - i. Determination of pH of soil.
 - ii. Total soluble salt
 - iii. Estimation of calcium, magnesium, phosphate, nitrate
- 4) Ion exchange:
 - i. Determination of exchange capacity of cation exchange resins and anion exchange resins.
 - ii. Separation of metal ions from their binary mixture.
 - iii. Separation of amino acids from organic acids by ion exchange chromatography

Section B: Organic Chemistry

- 1) Systematic qualitative analysis of organic compounds and preparation of suitable derivatives (including osazone formation in carbohydrates).
- 2) Separation of binary mixture of organic compounds using water, sodium bicarbonate, sodium hydroxide & hydrochloric acid.

Preparations

- 1) p-Nitroaniline from acetanilide
- 2) p-Bromoaniline from acetanilide

- 3) m-Nitroaniline from nitrobenzene
- 4) Synthesis of di-benzalacetone from benzaldehyde (Claisen-Schmidt condensation)
- 5) Synthesis of benzyl alcohol and benzoic acid from benzaldehyde (Cannizaro reaction)

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- 3) Christian, Gary D; *Analytical Chemistry*, 6th Ed. John Wiley & Sons, New York, 2004.
- 4) Harris, Daniel C: *Exploring Chemical Analysis*, Ed. New York, W.H. Freeman, 2001.
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- 6) Skoog, D.A. Holler F.J. & Nieman, T.A. *Principles of Instrumental Analysis*, Cengage Learning India Ed.
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- 8) Ditts, R.V. *Analytical Chemistry; Methods of Separation*, van Nostrand, 1974.
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		Day to day performance: 12 marks

B.	External Examination (100% syllabus)	Examination: 20 marks
		(Two practicals of 10 marks each)
		Viva-voce: 05 marks

SEMESTER-VI

(For examinations to be held in May 2019, 2020 & 2021)

Title: GREEN METHODS IN CHEMISTRY (Skill Enhancement Course SEC-4)

Course No.: UCHTS 601

Maximum Marks: 100

Credits: 04

External Examination: 80 marks

Time: 03 Hrs

Internal Assessment: 20 marks

Theory and Hand-on Experiments

Introduction: Definitions of Green Chemistry. Brief introduction of twelve principles of Green Chemistry, with examples, special emphasis on atom economy, reducing toxicity, green solvents, Green Chemistry and catalysis and alternative sources of energy, Green energy and sustainability

The following Real world Cases in Green Chemistry should be discussed:

- 1) Surfactants for carbon dioxide – Replacing smog producing and ozone depleting solvents with CO₂ for precision cleaning and dry cleaning of garments.
- 2) Designing of environmentally safe marine antifoulant.
- 3) Rightfit pigment: Synthetic azo pigments to replace toxic organic and inorganic pigments.
- 4) An efficient, green synthesis of a compostable and widely applicable plastic (poly lactic acid) made from corn.

Practicals

- 1) Preparation and characterization of biodiesel from vegetable oil.
- 2) Extraction of D-limonene from orange peel using liquid CO₂ prepared from dry ice.
- 3) Mechano chemical solvent free synthesis of azomethine.
- 4) Solvent free, microwave assisted one pot synthesis of phthalocyanine complex of copper(II).

Reference Books:

- 1) Anastas, P.T. & Warner, J.K. *Green Chemistry- Theory and Practical*, Oxford University Press (1998).
- 2) Matlack, A.S. *Introduction to Green Chemistry*, Marcel Dekker (2001).
- 3) Cann, M.C. & Connely, M.E. *Real-World cases in Green Chemistry*, American Chemical Society, Washington (2000).
- 4) Ryan, M.A. & Tinnesand, M. *Introduction to Green Chemistry*, American Chemical Society, Washington (2002).
- 5) Sharma, R.K.; Sidhwani, I.T. & Chaudhari, M.K. *Green Chemistry Experiments: A monograph* I.K. International Publishing House Pvt Ltd. New Delhi, Bangalore.
- 6) Lancaster, M. *Green Chemistry: An introductory text* RSC publishing, 2nd Edition.

- 7) Sidhwani, I.T., Saini, G., Chowdhury, S., Garg, D., Malovika, Garg, N. Wealth from waste: A green method to produce biodiesel from waste cooking oil and generation of useful products from waste further generated “A Social Awareness Project”, *Delhi University Journal of Undergraduate Research and Innovation*, **1(1)**: 2015.
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NOTE FOR PAPER SETTERS

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